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alphabetized both by title and by author.

## **Chapter 6 Chemical Bonds Answers**

CHAPTER 6 REVIEW

Chemical Bonding

SECTION 1 SHORT

ANSWER Answer the following questions in the space provided. 1. a A chemical bond between atoms results from the attraction between the valence electrons and of different atoms. (a)

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## Answers

nuclei (c) isotopes (b)  
inner electrons (d)  
Lewis structures 2. b A  
covalent bond consists  
of (a) a shared  
electron.

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Molecules. How many  
electrons should  
Helium have around its  
Lewis dot model?

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Answers

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Access Free Chapter 6  
Chemical Bonding  
Section 2 Covalent  
Answer Key malleable  
and ductile but ionic-  
crystalline compounds  
are not. The metallic  
bond is the same in all  
directions throughout  
the metallic structure  
allowing the atoms to  
slide past each other.

**Chapter 6 Chemical**  
*Page 7/27*

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## Answers

### **Bonding Section 2 Covalent Answer Key**

Chapter 6: Chemical Bonds Review Answer Key. What is the most stable group on the Periodic Table? Why is this group stable? Noble Gases (Group 18) because they have 8 valence electrons which means their outer energy level is full. Name the most reactive group(s) on the Periodic Table?



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Why are they reactive?

**Name** \_\_\_\_\_

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A chemical bond is the force that holds atoms or ions together as a unit. An ionic bond is the force that holds cations and anions together. An ionic bond forms when electrons are transferred from one atom to another.

Formation of Ionic

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bonds, Lewis dot  
structures, resonance  
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A Is it getting hot  
outside?

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ionic. A type of bond

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## Answers

that holds cations and anions together.

Covalent. A type of bond in which two atoms share a pair of valence electrons.

Molecule.

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## Answers

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CHAPTER 6 REVIEW

Chemical Bonding

SECTION 1 SHORT

ANSWER Answer the

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a A chemical bond

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different atoms. (a)

nuclei (c) isotopes (b)

inner electrons (d)

Lewis

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## **Chapter 6 Review Chemical Bonding Answers**

Chemical Bonding  
Class Date Section  
Quiz: Covalent Bonding  
and Molecular  
Compounds In the  
space provided, write  
the letter of the term  
or phrase that best  
completes each  
statement or best  
answers each question.  
1 If two covalently  
bonded atoms move

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closer than a distance of the bond length, the potential energy of the atoms a. becomes negative.

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607 Chapter 6, Lesson  
7: Energy Changes in  
Chemical Reactions.

Key Concepts • If two substances react and the temperature of the mixture decreases, the reaction is endothermic.

## **Chapter 6, Lesson 7: Energy Changes in Chemical Reactions**

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